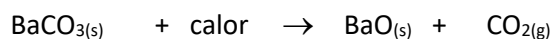
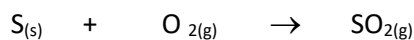
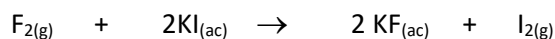
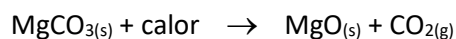
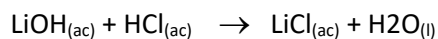
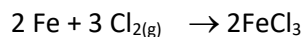
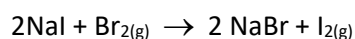
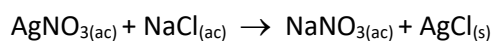
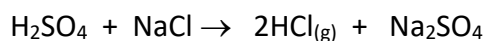
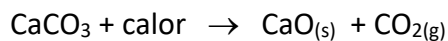
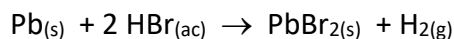
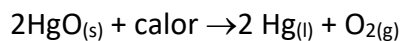
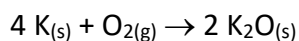
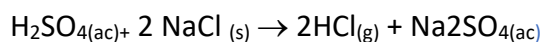
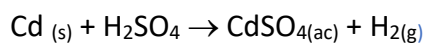


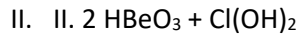
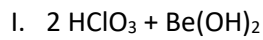
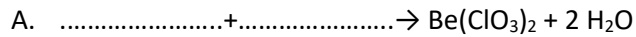
TRABAJO PRÁCTICO N°2: REACCIONES QUÍMICAS

1. Dadas las siguientes reacciones:

- Identifique las reacciones de doble desplazamiento.
- Identifique las reacciones de óxido-reducción.
- Identifique las reacciones de descomposición.
- Identifique las reacciones de combinación.
- Identifique las reacciones de desplazamiento.

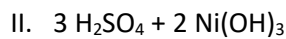
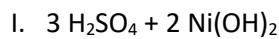
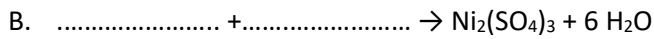


2. En las siguientes ecuaciones, elija los reactivos correspondientes (opción I o II) y luego elija el nombre de uno de los reactivos o productos (opción III o IV)



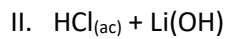
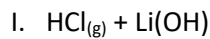
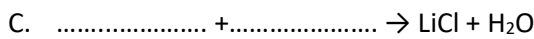
III. Clorato de berilio

IV. Clorito de berilio



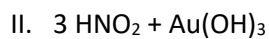
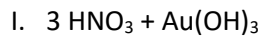
III. Sulfito níqueloso

IV. Sulfato níquelico



III. Hidróxido de litio

IV. Cloruro de hidrógeno



III. Ácido nitroso

IV. Nitrato aurioso

3. Dado los siguientes reactivos, decir si es factible o no la reacción entre ellos. En Caso afirmativo, escribir la ecuación química y balancear.

Tabla de potenciales de oxidación:

Potasio – Sodio- Aluminio- Cadmio- Cobalto- Níquel- Estaño- Plomo- Hidrogeno- Antimonio- Cobre -Mercurio -Plata -Platino

